organic papers

Acta Crystallographica Section E Structure Reports Online

ISSN 1600-5368

Chong-Gang Duan, Jiong Jia and Jian-Wu Wang*

School of Chemistry and Chemical Engineering, Shandong University, Jinan 250100, People's Republic of China

Correspondence e-mail: yugp2005@yahoo.com.cn

Key indicators

Single-crystal X-ray study T = 298 KMean $\sigma(C-C) = 0.005 \text{ Å}$ R factor = 0.054 wR factor = 0.145 Data-to-parameter ratio = 13.2

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

4-Amino-6-chloro-8-*p*-tolylpteridin-7(8*H*)-one dichloromethane hemisolvate

The title compound, $C_{13}H_{10}ClN_5O\cdot0.5CH_2Cl_2$, crystallizes with two independent 4-amino-6-chloro-8-*p*-tolylpteridin-7(8*H*)-one molecules and one dichloromethane molecule in the asymmetric unit. The bond lengths and angles in the molecules are within normal ranges. Intermolecular N-H···N hydrogen bonds link the two independent molecules into hydrogen-bonded dimers. The crystal packing is further stabilized by van der Waals forces.

Comment

Pteridine-like molecules are reported to be adenosine kinase inhibitors with modest potency (Perner *et al.*, 2003). In order to find novel compounds with high potency and high membrane permeability, several C6- and N8-substituted 4aminopteridin-7(8*H*)-one compounds have been designed. Compound 4-amino-6-chloro-8-*p*-tolylpteridin-7(8*H*)-one is one of the key intermediates. In this paper, we report the crystal structure of the title compound, (I).



Compound (I) crystallizes with two independent 4-amino-6chloro-8-*p*-tolylpteridin-7(8*H*)-one molecules (Fig. 1) and one dichloromethane molecule in the asymmetric unit. The bond lengths and angles in the molecules are within normal ranges (Allen *et al.*, 1987). Atoms Cl1, O1 and N5 lie in the mean plane (P1) of the C1–C6/N1–N4 7,8-dihydropteridine ring, the largest deviation being 0.062 (2) Å for atom N5. The dihedral angle between plane P1 and the C7–C12 benzene ring is 80.86 (3)°. Atoms Cl2, O2, N10 and C20 lie in the mean plane (P2) of the C14–C19/N6–N9 7,8-dihydropteridine ring, the largest deviation being 0.039 (4) Å for atom N10. The dihedral angle between plane P2 and the C20–C25 benzene ring is 71.67 (2)°.

Intermolecular $N-H\cdots N$ hydrogen bonds (Table 1) link two independent molecules into hydrogen-bonded dimers. The crystal packing (Fig. 2) is further stabilized by van der Waals forces. Received 23 May 2006 Accepted 31 May 2006

© 2006 International Union of Crystallography

All rights reserved



Figure 1

The asymmetric unit of (I), with displacement ellipsoids drawn at the 40% probability level. The solvent molecule has been omitted for clarity.

Experimental

4-Amino-8-*p*-tolylpteridine-6,7(5*H*,8*H*)-dione (2.69 g, 0.01 mol), phosphorus pentachloride (4.16 g, 0.02 mol) and phosphorus oxychloride (20 ml) were heated to reflux for 50 min. The mixture was concentrated to 10 ml *in vacuo*. After cooling to room temperature, chloroform (100 ml) was added. Water (100 ml) was then added dropwise while keeping the temperature below 293 K, and the mixture was stirred for 2 h. The organic layer was separated and the aqueous layer was extracted with chloroform. The combined organic layers were dried with anhydrous magnesium sulfate and concentrated *in vacuo*, and the residue was chromatographed on silica gel (ethyl acetate) to give 4-amino-6-chloro-8-*p*-tolylpteridin-7(8*H*)-one in dichloromethane at room temperature over a period of one week.

Crystal data

$C_{13}H_{10}ClN_5O{\cdot}0.5CH_2Cl_2$
$M_r = 330.18$
Monoclinic, $P2_1/c$
$a = 7.4644 (14) \text{\AA}$
b = 16.138 (3) Å
c = 24.402 (5) Å
$\beta = 97.676 \ (3)^{\circ}$
$V = 2913.1 (10) \text{ Å}^3$

Data collection

Bruker SMART CCD area-detector diffractometer φ and ω scans Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996) $T_{\min} = 0.899, T_{\max} = 0.935$ Z = 8 $D_x = 1.506 \text{ Mg m}^{-3}$ Mo K\alpha radiation $\mu = 0.45 \text{ mm}^{-1}$ T = 298 (2) KBlock, colourless $0.24 \times 0.18 \times 0.15 \text{ mm}$

14411 measured reflections 5123 independent reflections 3183 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.045$ $\theta_{\text{max}} = 25.0^{\circ}$



Figure 2

A partial packing diagram, showing the hydrogen-bonded (dashed lines) dimers.

Refinement

Refinement on F^2
$R[F^2 > 2\sigma(F^2)] = 0.054$
$vR(F^2) = 0.145$
S = 1.01
5123 reflections
389 parameters
H-atom parameters constrained

$$\begin{split} &w = 1/[\sigma^2(F_o^2) + (0.0672P)^2 \\ &+ 0.1625P] \\ &where \ P = (F_o^2 + 2F_c^2)/3 \\ (\Delta/\sigma)_{max} < 0.001 \\ \Delta\rho_{max} = 0.31 \ e \ \text{\AA}^{-3} \\ \Delta\rho_{min} = -0.32 \ e \ \text{\AA}^{-3} \end{split}$$

Table 1

Hydrogen-bond	geometry	(Å,	°).	•
---------------	----------	-----	-----	---

$D - H \cdots A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdot \cdot \cdot A$
N5−H5 <i>B</i> ···N8 ⁱ	0.86	2.08	2.937 (4)	179
$10 - H10A \cdots N3^n$	0.86	2.23	3.089 (4)	176
Symmetry codes: (i) $-x$	$-1, y - \frac{1}{2}, -z$	$+\frac{1}{2}$; (ii) $-x - 1$	$y + \frac{1}{2}, -z + \frac{1}{2}$	170

All H atoms were placed in calculated positions, with C-H = 0.93 or 0.96 Å and N-H = 0.86 Å, and included in the final cycles of refinement using a riding model, with $U_{iso}(H) = 1.2U_{eq}(C)$ for the aryl, methine and N-bound H atoms, and $1.5U_{eq}(C)$ for the methyl H atoms.

Data collection: *SMART* (Bruker, 1998); cell refinement: *SAINT* (Bruker, 1999); data reduction: *SAINT*; program(s) used to solve structure: *SHELXTL* (Bruker, 1999); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

References

Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). J. Chem. Soc. Perkin Trans. 2, pp. S1–19.

- Bruker (1998). SMART. Version 5.054. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (1999). SAINT (Version 6.36a) and SHELXTL (Version 5.10). Bruker AXS Inc., Madison, Wisconsin, USA.
- Perner, R. J., Gu, Y.-G., Lee, C.-H., Bayburt, E. K., McKie, J., Alexander, K. M., Kohlhaas, K. L., Wismer, C. T., Mikusa, J., Jarvis, M. F., Kowaluk, E. A. & Bhagwat, S. S. (2003). J. Med. Chem. 46, 5249–5257.
- Sheldrick, G. M. (1996). SADABS. Version 2.0. University of Göttingen, Germany.